

- Find the empirical formula of a compound containing: 92.31 % C, and 7.69 % H.
- Find the empirical formula of a compound containing: 25.92 % N, and 74.07 % O.
- Find the empirical formula of a compound containing: 43.67 % P, and 56.33 % O.
- Find the empirical formula of a compound containing: 19.32 % Ca, 34.30 % Cl, and 46.38 % O.
- Find the empirical formula of a compound containing: 64.86 % C, 13.52 % H, and 21.62 % O.
- Find the empirical formula of a compound containing: 3.06 % H, 31.63 % P, and 65.30 % O.
- Find the empirical formula of a compound containing: 24.56 % K, 34.81 % Mn, and 40.50 % O.
- Find the empirical formula of a compound containing: 37.70 % Na, 22.95 % Si, and 39.34 % O.
- Find the empirical formula of a compound containing: 31.25 % Ca, 18.75 % C, and 50.00 % O.
- Find the empirical formula of a compound containing: 52.17 % C, 13.04 % H, and 34.78 % O.
- Find the empirical formula of a compound containing:
28.05 % Na, 29.26 % C, 3.66 % H, and 39.02 % O.
- Find the empirical formula of a compound containing:
21.21 % N, 6.06 % H, 24.24% S, and 48.48 % O.
- Find the empirical formula of a compound containing:
24.69 % Ca, 1.23 % H, 14.81 % C, and 59.26 % O.
- What is the percentage composition of each element of the following:
a) CH₂Cl₂ b) C₉H₁₀O₄
- Find the empirical formula of a compound containing:
a) 37.5 % C, 12.5 % H, and 50.5 % O b) 19.9 % C, 0.4 % H, 13.3 % O, and 66.4 % Br
- A hydrocarbon is burned and releases 71.0 % carbon dioxide and 29.0 % water vapor.
What is the empirical formula of the hydrocarbon?
- Another hydrocarbon is burned and releases water vapor and 78.6 % carbon dioxide.
What is the empirical formula of this hydrocarbon?
- A carbohydrate sample has a mass of 100. g. It is totally burned, producing 76.5 g of CO₂ and
23.5 g of H₂O. What is the empirical formula?